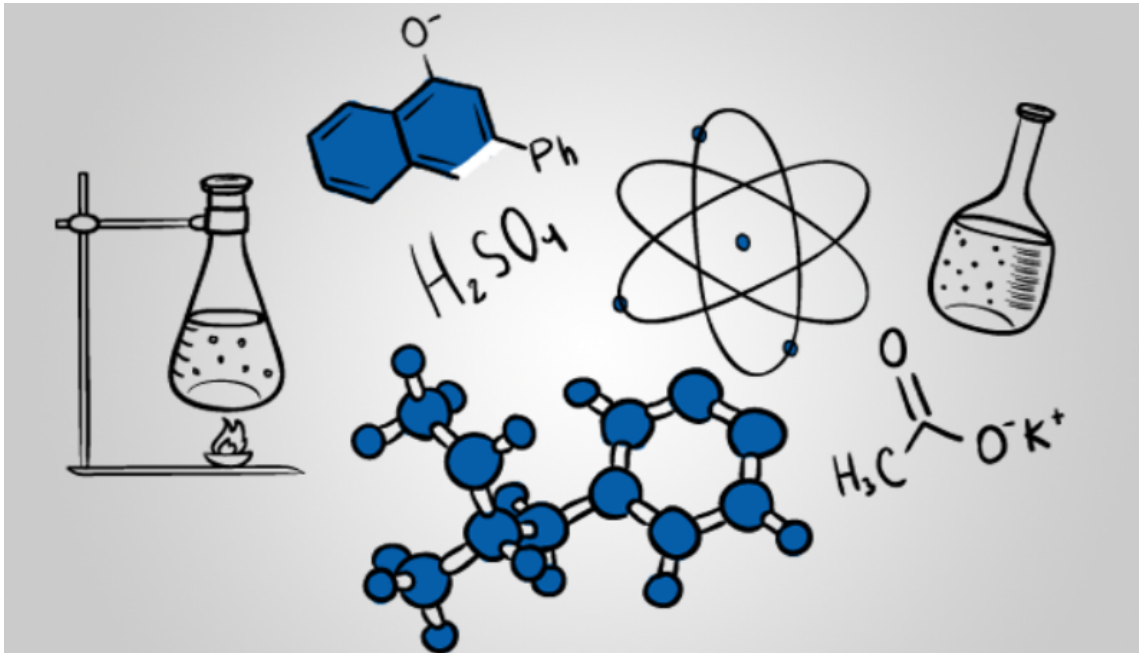


Welcome! Please grab your ISN and have a seat!
Please complete the element review in your classroom!!!



Oct 31-10:49 AM

Strand **5** TITLE: **Compounds**

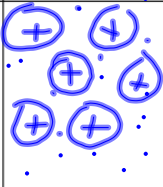
Page #	Page Title
49	Strand 5 TDC
50	Strand 5 Calendar
51-52	Strand 5 WWK
53-54	Types of Bonds

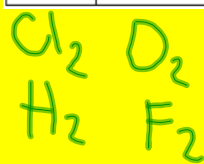
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W&W (pg 51)

1. compound - pure substance where 2 or more elements are chemically bonded.
2. molecule - 2 or more elements or compounds covalently bonded together.
3. mixture - 2 or more elements or compounds that coexist but are not chemically bonded.
4. salt - compound formed by one cation and one anion
5. electrolyte - substance that conducts electricity when dissolved in water, but not as a solid

Oct 31-11:43 AM

	Ionic	Covalent	Metallic
What bonds?	Metal (+) nonmetal (-)	2 nonmetals	2 metals
What happens?	metal loses e^- nonmetal gains e^-	share e^- pairs to become stable	atoms share e^- via sea of electrons
Properties	<ul style="list-style-type: none"> • strong bonds • high melting & boiling point • electrolytes • hard & brittle 	<ul style="list-style-type: none"> • weak bonds • low melting & boiling pts • nonconductors • malleable 	<ul style="list-style-type: none"> • flexible • great conductors • lattice formation
example	$Mg \overset{+}{\curvearrowright} \overset{-}{O}$ $Mg^{2+} \quad O^{2-}$	$Cl \overset{\cdot\cdot}{\curvearrowright} Cl$ $Cl:Cl$ (diatomic)	



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- 1
- 2
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- 11

COMPOUND	BOND TYPE	LEWIS ELECTRON DOT
NaCl		
MgCl ₂	I	$\text{Mg}^{2+} \cdot \text{Cl}^- \cdot \text{Cl}^-$
Br ₂		
N ₂		
HCl		
H ₂ O		
CO ₂		
NH ₃		
CH ₄		
4Ti		
3Ag		

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Oct 31-11:49 AM