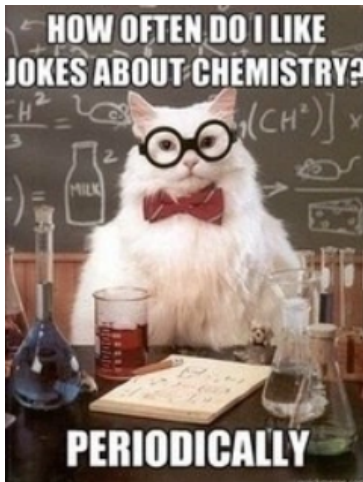


1. What is the density of CO gas if 0.196 g occupies a volume of 100 ml?

Answer _____

2. A block of wood 3 cm on each side has a mass of 27 g. What is the density of the block? (Hint, don't forget to find the volume of the wood.)



Answer _____

Aug 22-10:10 AM

3. An irregularly shaped stone was lowered into a graduated cylinder holding a volume of water equal to 2 ml. The height of the water rose to 7 ml. If the mass of the stone was 25 g, what was its density?

Answer _____

4. A 10.0 cm^3 sample of copper has a mass of 89.6 g. What is the density of copper?

Aug 22-11:58 AM

5. Silver has a density of 10.5 grams/cm^3 and gold has a density of 19.3 g/cm^3 . Which would have the greater mass, 5cm^3 of silver or 5cm^3 of gold?

Answer _____

6. Five mL of ethanol has a mass of 3.9 g, and 5.0 mL of benzene has a mass of 4.4 g. Which liquid is denser?

Answer _____

Aug 22-11:59 AM

7. A sample of iron has the same dimensions of 2 cm x 3 cm x 2 cm. If the mass of this rectangular-shaped object is 94 g, what is the density of iron?

Aug 22-12:00 PM

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Aug 22-10:10 AM

textbook:

OMK6-LD4L-OVQO

learnsmart:

TZ7P - 4KRS - Z920

Aug 23-10:46 AM

WWK

5. accuracy - how close your observation is to accepted value (right answer)

$$A = 5.2g$$

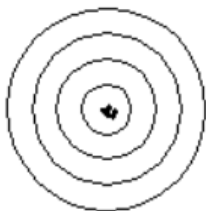
4.8g
5.4g
5.6g

6. precision - how close together a set of observations is.

1.1mL 1.17mL
1.2mL 1.08mL

3mL 1.4mL 2mL
3.2mL

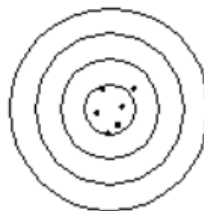
Aug 23-10:09 AM



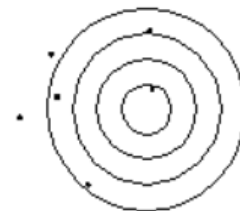
accurate and precise



precise but not accurate



accurate but not precise



not accurate or precise

Aug 23-1:42 PM

TOC 19-20 Scientific notation and significant figures

7 sig figs.

SIGNIFICANT

- all non zero #'s are significant
- any zero that is b/w 2 nonzeros.
- any trailing zeroes after a decimal.

0.0001740600

NOT SIGNIFICANT

- Lonely zeroes before a decimal
- leading zeroes after a decimal

Aug 22-12:00 PM

TOC 19-20 Scientific notation and significant figures
move decimal until the # is b/w 1-10*

SCIENTIFIC NOTATION

0.3.6.4.0

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RIGHT (-)
LEFT (+)

3.640×10^{-3}

Aug 22-12:02 PM

TOC 19-20 Scientific notation and significant figures

Regular Notation	Scientific Notation
420,000.	
735,000,000.	
.00897	
.0000014	

Aug 22-12:09 PM

1. Indicate how many significant figures there are in each of the following measured values. Then, write the number in scientific notation.

	Sig. figs.	sci note
246.32	_____	_____
107.854	_____	_____
100.3	_____	_____
0.678	_____	_____
1.008	4	1.008×10^0
0.00340	_____	_____
14,600	_____	_____
0.0001	_____	_____
700000	_____	_____
350,670	_____	_____
1,0000	_____	_____
320001	_____	_____

**examples
page 19**

Aug 22-12:09 PM