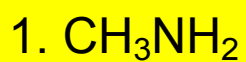
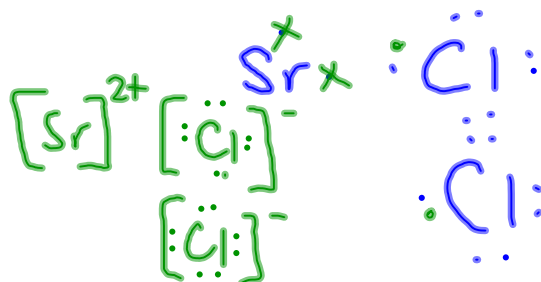
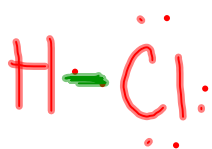
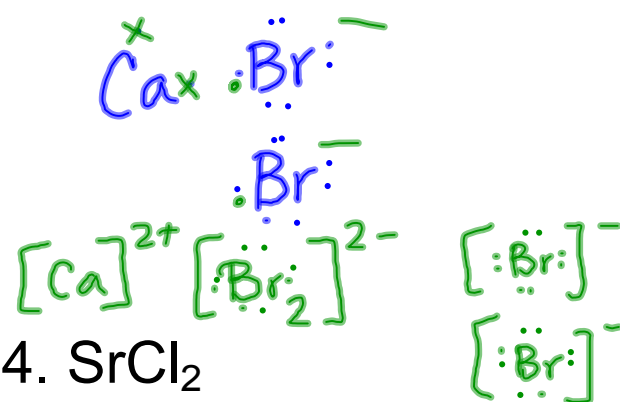
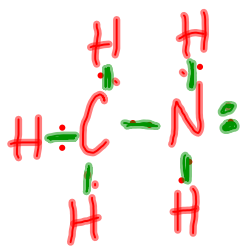
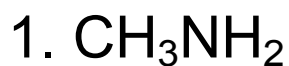


Welcome! Please draw the Lewis dot structures at the bottom of page 53 in your ISM!!!!



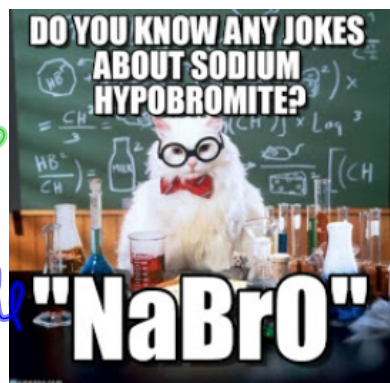
Nov 2-10:25 AM



Nov 2-11:54 AM

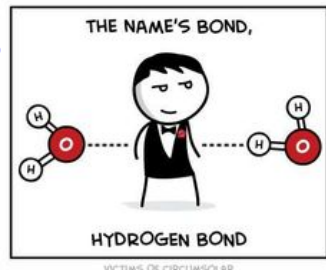
- 1) Na_2CO_3 sodium carbonate
- 2) NaOH sodium hydroxide
- 3) MgBr_2 magnesium bromide
- 4) KCl potassium chloride
- 5) FeCl_2 Iron(II) chloride
- 6) FeCl_3 Iron(III) chloride
- 7) Zn(OH)_2 zinc(II) hydroxide
- 8) BeSO_4 beryllium sulfate
- 9) CrF_2 Chromium(II) fluoride
- 10) Al_2S_3 aluminum sulfide

WELCOME!! GET OUT YOUR HOMEWORK, AND TRY TO WORK ON 11-15!!!



Nov 2-11:03 AM

- 22) sodium phosphide ~~$\text{Na}^+ \text{P}^{3-}$~~ Na_3P
- 23) magnesium nitrate ~~$\text{Mg}^{2+} (\text{NO}_3)^-$~~ $\text{Mg}(\text{NO}_3)_2$
- 24) lead (II) sulfite ~~$\text{Pb}^{2+} (\text{SO}_3)^{2-}$~~ PbSO_3
- 25) calcium (phosphate) ~~$\text{Ca}^{2+} \text{PO}_4^{3-}$~~ $\text{Ca}_3(\text{PO}_4)_2$
- 26) ammonium (sulfate) ~~$\text{NH}_4^+ \text{SO}_4^{2-}$~~ $(\text{NH}_4)_2\text{SO}_4$
- 27) silver cyanide ~~AgCN~~
- 28) aluminum sulfide _____
- 29) beryllium chloride ~~$\text{Cu}^+ \text{As}^{3-}$~~ Cu_3As
- 30) copper (I) arsenide _____
- 31) iron (III) oxide _____



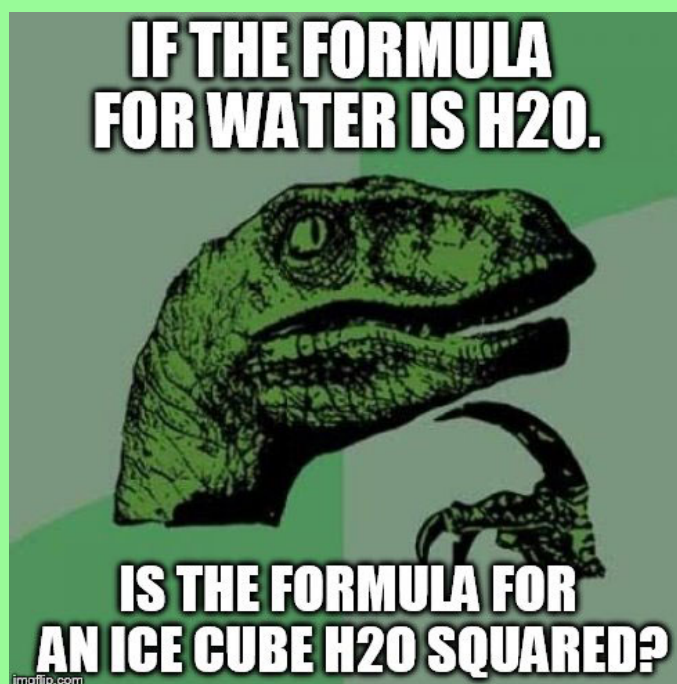
Nov 2-11:04 AM

- 32) gallium nitride _____
- 33) iron (II) bromide _____
- 34) vanadium (V) phosphate _____
- 35) calcium oxide _____
- 36) magnesium acetate _____
- 37) aluminum sulfate _____
- 38) copper (I) carbonate _____
- 39) barium oxide _____
- 40) ammonium sulfite _____
- 41) silver bromide _____
- 42) lead (IV) nitrite _____



Nov 2-11:06 AM

Welcome! Please watch the EdPuzzle on naming acid compounds in your Google classroom!



Nov 2-11:32 AM

WWK

6. acid- compounds that donate H^+ cations when dissolved in water.

Nov 14-8:20 AM

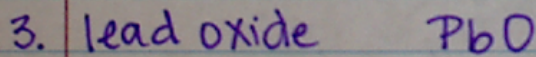
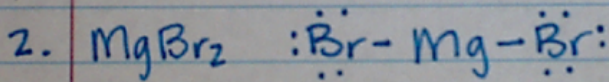
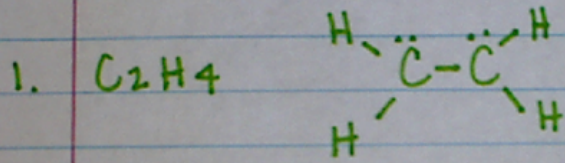
TOC PG 55-56 NAMING COMPOUNDS

CH ₄ How to Name Covalent Compounds	
<p>From a formula...</p> <p>prefix-element prefix-element-ide</p> <p>*If the 1st element is 1, do not start w/mona</p>	<p>From a name⁵</p> <p>phosphorus pentachloride PCl₅ dinitrogen trifluoride N₂F₃</p>
CO ₂ H ₂ O How to Name Acids	
<p>From a formula...</p> <p>• binary - hydro- el- ic acid • oxyacids - ION- ic acid</p>	<p>From a name...</p> <p>chromic acid hydrofluoric acid</p> <p>$H^+ \leftarrow \overset{2-}{\text{CrO}_4}$ $H^+ \leftarrow \overset{-}{\text{F}}$</p> <p>H₂CrO₄ HF</p>

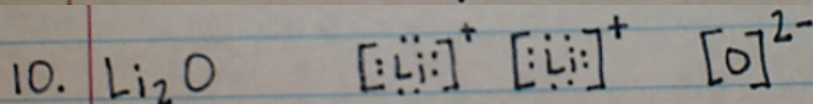
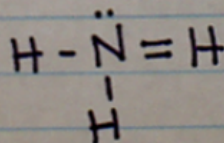
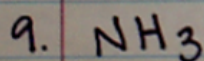
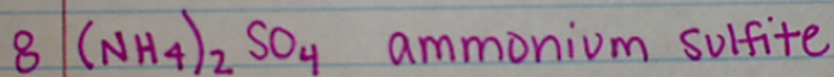
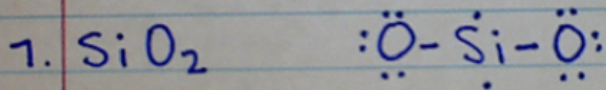
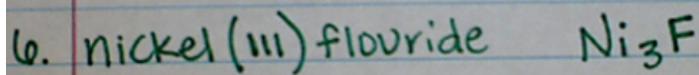
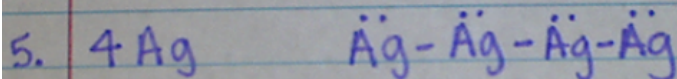
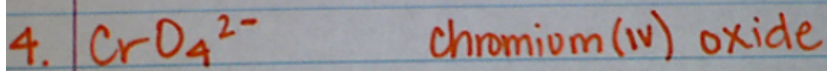
Nov 2-10:33 AM

What is WRONG here?

* Complete sentence
stating what's
wrong. Then, show
the correct
answer!



Nov 8-11:54 AM



Nov 8-11:54 AM

NAME:

* Identify the type of compound (covalent or acid), then name it!

1. P_4S_5 _____
2. H_3PO_4 _____
3. SF_6 _____
4. CH_4 _____
5. HBr _____
6. N_2O_5 _____
7. HCl _____
8. H_2SO_4 _____
9. Si_2Br_6 _____
10. HF _____

Nov 14-12:35 PM

* Identify the type of compound (C or A), then write the formula!

11. diphosphorus pentoxide _____
12. hydroiodic acid _____
13. iodine pentafluoride _____
14. nitric acid _____
15. phosphoric acid _____
16. methane _____
17. dinitrogen trioxide _____
18. sulfuric acid _____
19. carbon tetrachloride _____
20. hydrosulfuric acid _____

Nov 14-12:36 PM