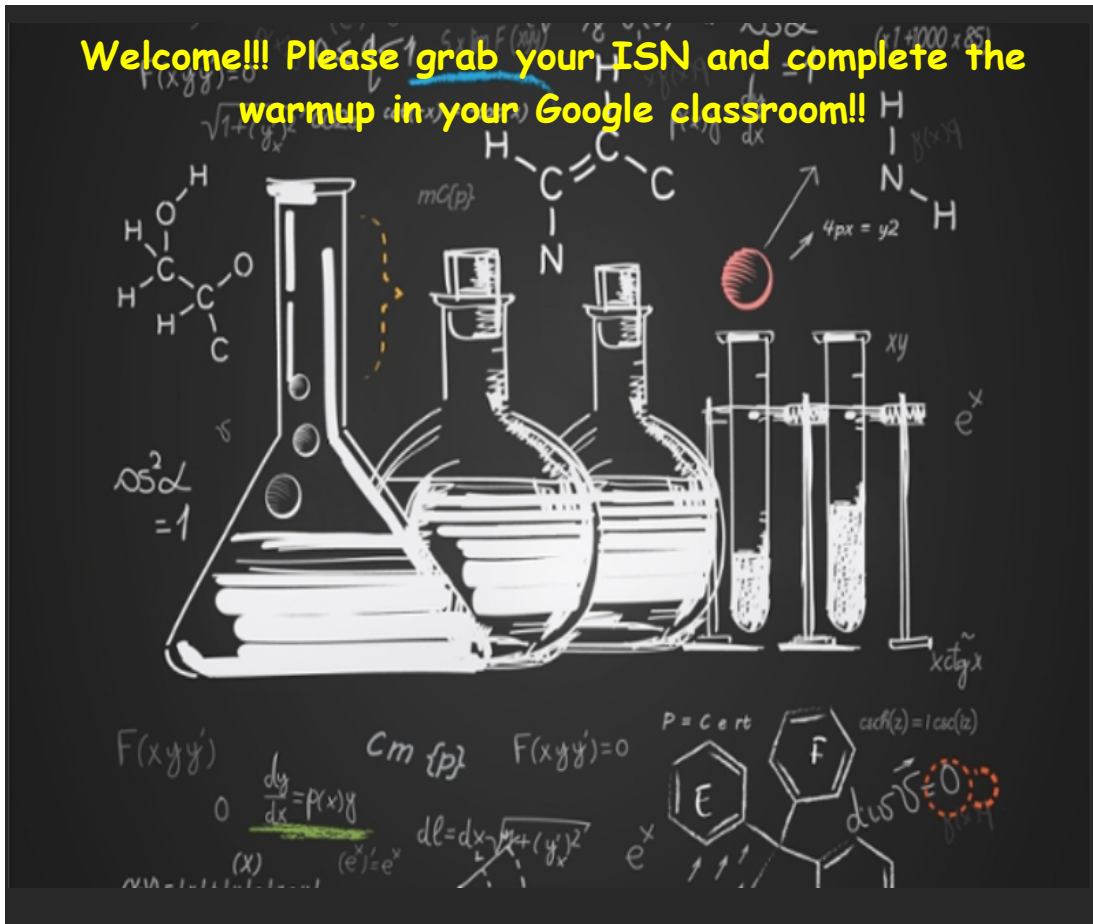
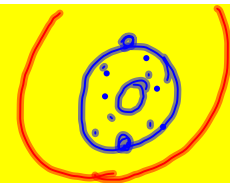


Welcome!!! Please grab your ISN and complete the warmup in your Google classroom!!



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WWWK (PG. 33)



14. electron configuration-written notation that describes the energy and location of electrons by using quantum numbers.

15. ground state-electrons having the lowest possible energy state.

16- excited state-electron gains energy and moves up an energy level

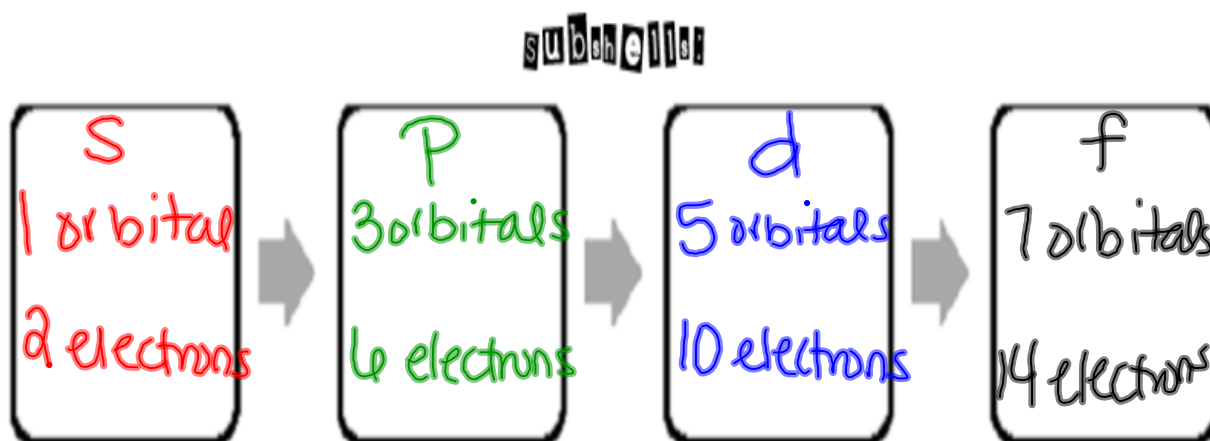
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TOC pg 45-46 Electron Configuration

1. Electron arrangement determines chemical behavior & bonding.
2. The electron shells that surround the nucleus can be broken down into subshells.
3. Each subshell can be broken down into individual orbitals.
4. The number of orbitals a subshell has determines the energy level of the subshell.

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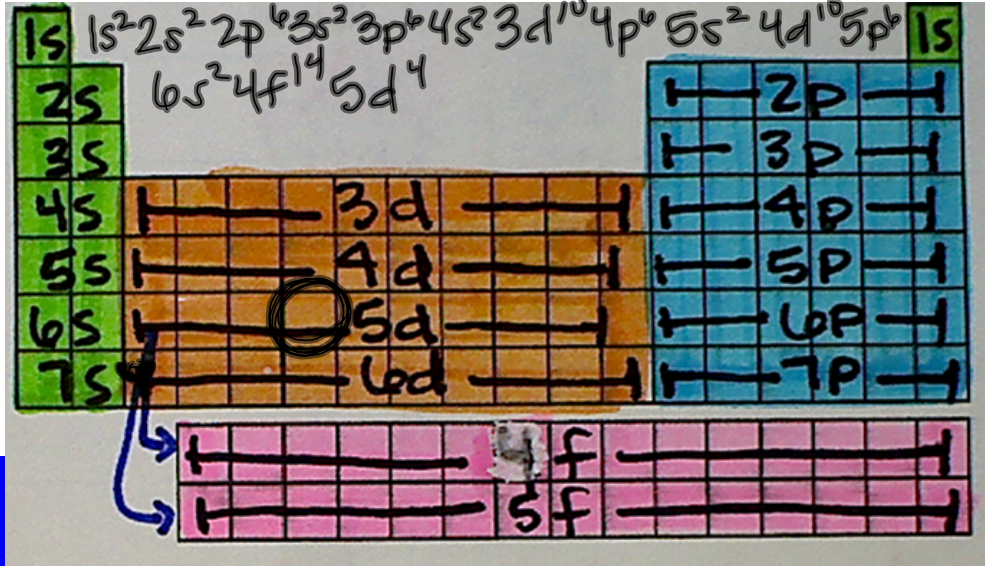
TOC pg 45-46 Electron Configuration



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TOC pg 45-46 Electron Configuration

Periodically, energy levels fill in a horizontal direction.



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TOC pg 45-46 Electron Configuration

Electron Configuration Rules:

Aufbau Principle	Hund's Rules	Pauli's Exclusion Principle

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pg 46 examples under graphic organizer

Aluminum

Iron

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pg 45 examples

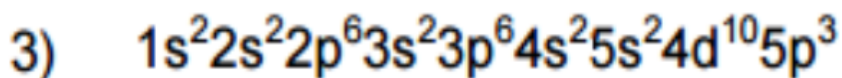
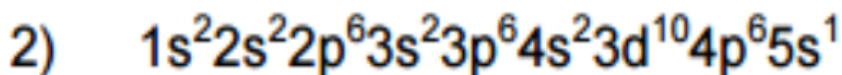
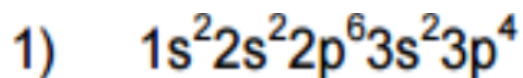
In the space below, write the electron configurations of the following elements:

- 1) sodium
- 2) bromine
- 3) barium
- 4) neptunium


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pg 45 examples

Determine what elements are denoted by the following electron configurations:



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Homework

Chemistry I Practice - "Electron Configurations"

Use the following electron configurations and your periodic table to identify the element:

1. $1s^2 2s^2 2p^6 3s^2 3p^5$ 2. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$ 3. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^1$

4. Describe the method that you used to solve problems 1 - 3. Be specific.

Use the following clues to identify the element. Show any figuring in the space below.

5. This element has a 3p sublevel that contains 3 electrons.

6. This element has a 4s sublevel with 2 electrons for its outermost electrons.

7. This element has 1 electron in its 3d sublevel.

8. This element has 5 electrons in its 5p sublevel

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