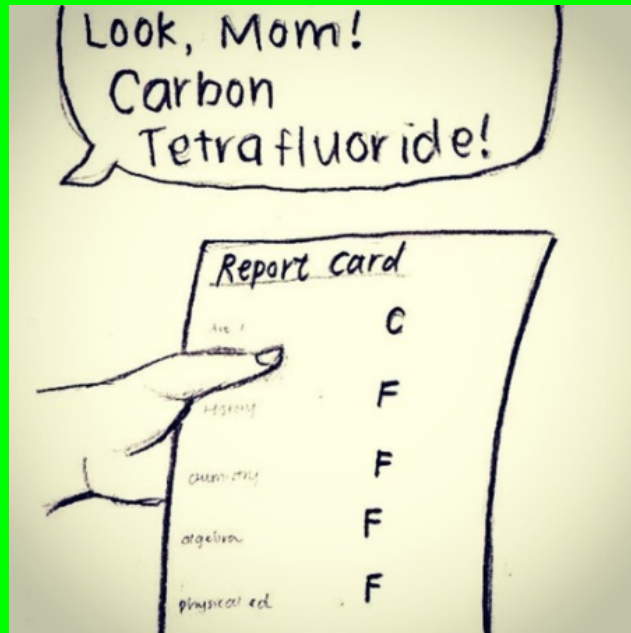


Complete the naming game in the google classroom.



Nov 15-11:40 AM

*** There are 3 types of bonds: ***



*** Identify the type of bond each compound will make: (I) ionic (C) covalent (M) metallic ***

- | | |
|-------------------------------|-----------------------------|
| 1. <u>I</u> NaBr | 4. <u>C</u> CH ₄ |
| 2. <u>I</u> CaCl ₂ | 5. <u>M</u> 3Au |
| 3. <u>M</u> 4Ti | 6. <u>C</u> CO ₂ |

Nov 15-7:51 AM

** Fill in the vocab blanks: ***

diatomic cation polyatomic ion
 stable transition metal covalent
 metallic ionic acid

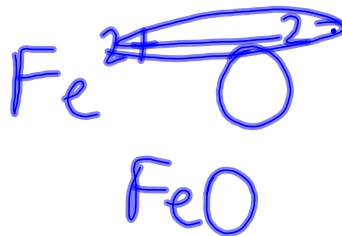
7. Transition metals need Roman numerals to indicate charges in ionic compounds.
8. In an ionic compound, the cation is named first.
9. Two atoms of the same element bonded together covalently are called diatomic. N_2 O_2 F_2 Cl_2
10. Electrons are shared in a covalent bond.
11. When an element loses electrons, it forms a cation.
12. A compound that donates H^+ cations when dissolved in water is a (n) acid.
13. Elements bond together to become more stable.

Nov 15-7:51 AM

** Naming IONIC compounds. ***

14. Name the metal cation first, then the non metal anion.
15. K_2O potassium oxide
16. $(NH_4)_2S$ ammonium sulfide
17. $Mg(NO_3)_2$ magnesium nitrate
18. Iron (II) oxide FeO
19. lithium phosphate Li_3PO_4
20. aluminum hydroxide $Al(OH)_3$

magnesium
nitrate



Nov 15-7:52 AM

*** Naming COVALENT compounds ***

21. List the following prefixes:

2 - di 5 - penta 7 - hepta 6 - hexa
 4 - tetra 9 - nona 3 - tri 1 - mono

22. methane CH_4 25. SeF_2 selenium difluoride

23. pentaboron silicide B_5Si 26. B_2Si diboron silicide

24. antimony tribromide $SbBr_3$ 27. P_4S_5 →
 tetraphosphorus
 pentasulfide

Nov 15-7:52 AM

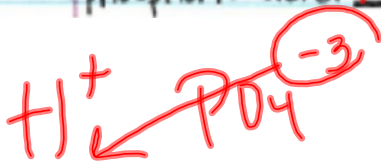


*** Naming ACIDS ***

28. Bromic acid $HBrO_3$ 31. HF hydrofluoric acid

29. hydrofluoric acid HF 32. H_2SO_4 sulfuric acid

30. phosphoric acid H_3PO_4 33. HBr hydrobromic acid



Nov 15-7:53 AM

*** Draw the Lewis Structure: ***

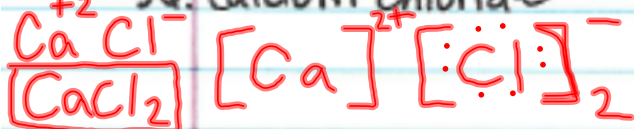
34. O^{2-}



35. Ge



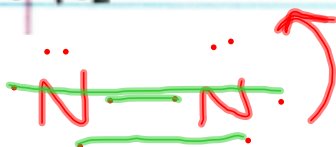
36. Calcium chloride



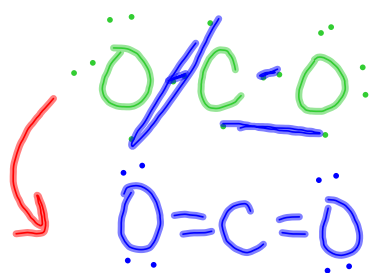
37. C_2H_2



38. N_2



39. CO_2



Nov 15-7:53 AM

BO_3^{3-} borate ion

BrO_3^- bromate ion

$C_2H_3O_2^-$ acetate ion (CH_3COO^-)

ClO_3^- chlorate ion

CN^- cyanide ion

CO_3^{2-} carbonate ion

CrO_4^{2-} chromate ion

IO_3^- iodate ion

MnO_4^- permanganate ion

NH_4^+ ammonium ion

NO_3^- nitrate ion

SeO_4^{2-} selenate ion

SiO_4^{4-} silicate ion

OH^- hydroxide ion

PO_4^{3-} phosphate ion

SO_4^{2-} sulfate ion

Nov 15-1:46 PM