

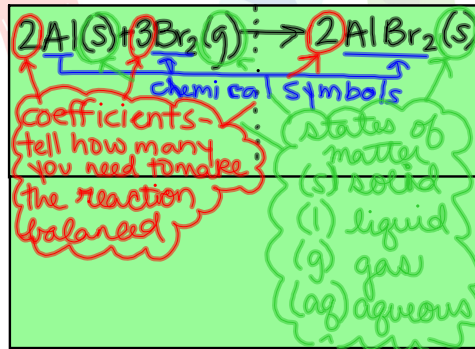
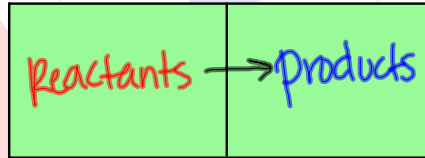
Welcome! Please grab your ISN
and have a seat!
Complete the game in your
google classroom!!

Nov 28-7:45 AM

WWK:

7. chemical reaction- atoms of one or more substances chemically rearrange to form new substances.
8. reactants- substances that you start with in a reaction
9. products- new substances formed during a reaction
10. coefficient- number in front of a compound telling how many of the compound reacts or is formed

Nov 28-8:43 AM



**** Always Diatomic ****

Oxygen Hydrogen Nitrogen Bromine Iodine Fluorine
Chlorine

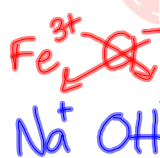
Nov 28-8:44 AM

How to Chemical Equations!

* You must have the same number of atoms of each element on both sides of the equation!!

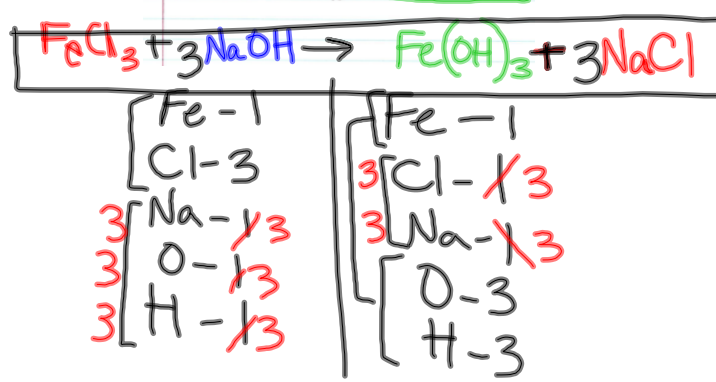
#lawofconservationofmass #knowurchemistry
 #doevenstudytoo #teachershocktagtoo

STEPS:	Example								
1. Write equation	$N_2 + H_2 \rightarrow NH_3$								
2. Count the # atoms of reactants/products	<table border="0"> <tr> <td>N = 2</td> <td>H = 2</td> <td>N = 1</td> <td>H = 3</td> </tr> <tr> <td>3N = 6</td> <td>2H = 6</td> <td>2N = 2</td> <td>6H = 6</td> </tr> </table>	N = 2	H = 2	N = 1	H = 3	3N = 6	2H = 6	2N = 2	6H = 6
N = 2	H = 2	N = 1	H = 3						
3N = 6	2H = 6	2N = 2	6H = 6						
3. Add coefficients	$N_2 + 3H_2 \rightarrow 2NH_3$								
molecules until the # atoms is the same on both sides!	* check! ✓								
* Typically, save O and H for last!	N = 2 H = 6 N = 2 H = 6								



Give it a TRY!

In water, iron (III) chloride reacts with sodium hydroxide, producing a solid iron (III) hydroxide and sodium chloride.



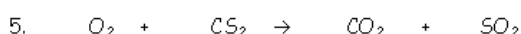
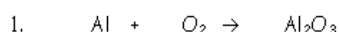
Nov 28-8:53 AM

Fill in the blanks with the most appropriate term:

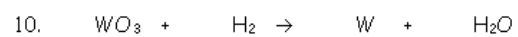
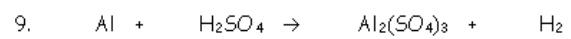
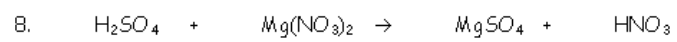
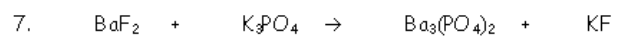
A chemical equation tells the story of a chemical reaction. Reactants are the starting substances in the reaction while products are the new substances that are formed. The large numbers in front of some of the formulas are called coefficients. These numbers are used to balance the equation because chemical reactions must obey the Law of conservation of Matter. The number of atoms of each element on both sides of the equation must be equivalent because matter cannot be created or destroyed. When balancing equations, the only numbers that can be changed are coefficients; remember that subscripts must never be changed in order to balance an equation.

Nov 28-8:57 AM

II. Balance the following equations:



Nov 28-8:58 AM



Nov 28-8:59 AM